

Staff Pre-Planning Template

Content: Chemistry (Brandalise) Grade: 10_____

Standards	Objectives
Week 1	
<p>HS-PS1-7. Use mathematical representations and provide experimental evidence to support the claim that atoms, and therefore mass, are conserved during a chemical reaction. Use the mole concept and proportional relationships to evaluate the quantities (masses or moles) of specific reactants needed in order to obtain a specific amount of product.</p> <p>HS-PS2-8(MA). Use kinetic molecular theory to compare the strengths of electrostatic forces and the prevalence of interactions that occur between molecules in solids, liquids, and gases. Use the combined gas law to determine changes in pressure, volume, and temperature in gases. [</p>	<p>SWBAT apply the Kinetic Molecular Theory when describing the properties of solids liquids and gases.</p>
Week 2	
<p>HS-PS1-7. Use mathematical representations and provide experimental evidence to support the claim that atoms, and therefore mass, are conserved during a chemical reaction. Use the mole concept and proportional relationships to evaluate the quantities (masses or moles) of specific reactants needed in order to obtain a specific amount of product.</p> <p>HS-PS2-8(MA). Use kinetic molecular theory to compare the strengths of electrostatic forces and the prevalence of interactions that occur between molecules in solids, liquids, and gases. Use the combined gas law to determine changes in pressure, volume, and temperature in gases. [</p>	<p>SWBAT apply the Kinetic Molecular Theory when describing the properties of solids liquids and gases.</p>
Week 3	
<p>HS-PS1-7. Use mathematical representations and provide experimental evidence to support the claim that atoms, and therefore mass, are conserved during a chemical reaction. Use the mole concept and proportional relationships to evaluate the quantities (masses or moles) of</p>	<p>SWBAT use the Kinetic Molecular Theory to describe verbally and mathematically the following relationships: pressure and volume, volume and temperature, pressure and</p>

<p>specific reactants needed in order to obtain a specific amount of product.</p> <p>HS-PS2-8(MA). Use kinetic molecular theory to compare the strengths of electrostatic forces and the prevalence of interactions that occur between molecules in solids, liquids, and gases. Use the combined gas law to determine changes in pressure, volume, and temperature in gases. [</p>	<p>temperature. (Boyle's Law, Charles' Law, and Gay-Lussac's Law)</p>
Week 4	
<p>HS-PS1-7. Use mathematical representations and provide experimental evidence to support the claim that atoms, and therefore mass, are conserved during a chemical reaction. Use the mole concept and proportional relationships to evaluate the quantities (masses or moles) of specific reactants needed in order to obtain a specific amount of product.</p> <p>HS-PS2-8(MA). Use kinetic molecular theory to compare the strengths of electrostatic forces and the prevalence of interactions that occur between molecules in solids, liquids, and gases. Use the combined gas law to determine changes in pressure, volume, and temperature in gases. [</p>	<p>SWBAT use the Kinetic Molecular Theory to describe verbally and mathematically the following relationships: pressure and volume, volume and temperature, pressure and temperature. (Boyle's Law, Charles' Law, and Gay-Lussac's Law)</p>

1. Include Topic/Concept/Theme on Student Work Packets. (*see template*)

2. Upload to School Website:

- Student Weekly Work Packets

3. Make Copy of Each and Put in Crate in Main Office of:

- Staff Weekly Planning Template (1 for each week total of 4 in Crate)
- Student Work Packets (1 for each week total of 4 in Crate)

Staff Weekly Work Template

Week (circle one) 1 2 3 4

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Topic/Theme: States of Matter and the Kinetic Molecular Theory

Concept:

Properties of solids, liquids, and gases and the Relationship to Temperature

Brief Introduction:

The properties of gases, liquids, and solids are described in terms of molecular motion. The effect of temperature on molecular motion as it relates to the various states of matter, in particular gases, is described.

1. Eureka Program 1 Inertia <https://www.youtube.com/watch?v=HRq-v4Gmzxc&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=1>
2. Eureka Program 2 Mass <https://www.youtube.com/watch?v=o5mL2Y2WNDs&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=2>
3. Eureka Program 7 Mass vs. Weight <https://www.youtube.com/watch?v=xfGll1REu2A&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=7>
4. Eureka Program 16 Solids <https://www.youtube.com/watch?v=4TPV3V39PMI&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=16>
5. Eureka Program 17 Liquids <https://www.youtube.com/watch?v=TxpwwPOwR30&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=17>
6. Eureka Program 18 Evaporation and Condensation https://www.youtube.com/watch?v=_QH52gQy0Ac&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=18

Other

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1. Eureka Program 19 Expansion and Contraction

<https://www.youtube.com/watch?v=KnyoamfPHY&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=19>

2. Eureka Program 20 Measuring Temperature

https://www.youtube.com/watch?v=aBDnIA_0Ru4&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=20

3. Eureka Program 21 Temperature vs. Heat

<https://www.youtube.com/watch?v=jb5Tv3G1vSc&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=21>

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https://www.youtube.com/watch?v=aBDnIA_0Ru4&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=20
3. Eureka Program 21 Temperature vs. Heat
<https://www.youtube.com/watch?v=jb5Tv3G1vSc&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=21>
4. Eureka Program 22 Atoms
<https://www.youtube.com/watch?v=nowdQow7ZH8&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=22>
5. Eureka 23 Electrons <https://www.youtube.com/watch?v=y-4W95dHVXc&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=23>
6. Eureka 24 Conduction <https://www.youtube.com/watch?v=Yitiw6Y7xZg&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=24>

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1. Eureka Program 22 Atoms <https://www.youtube.com/watch?v=nowdQow7ZH8&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=22>
2. Eureka 23 Electrons <https://www.youtube.com/watch?v=y-4W95dHVXc&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=23>
3. Eureka 24 Conduction <https://www.youtube.com/watch?v=Yitiw6Y7xZg&list=PL1o41VZtwLyVunxGUF-ZRkDTjNU01AGlz&index=24>

Other

****PDF Attached or Downloaded****

